

Surface Chemistry Ncert Solutions

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Solution: When a gas is adsorbed on the surface of a solid, its entropy decreases, i.e., ΔS is negative. Now, $\Delta G = \Delta H - T\Delta S$ For a process to be spontaneous, ΔG must be negative. As here, ΔS is negative, therefore, $T\Delta S$ is positive ΔG can be negative only if ΔH is negative Hence, adsorption is always exothermic. 5.9.

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we come across in surface chemistry may be pure compounds or solutions. The interface is normally a few molecules thick but its area depends on the size of the particles of bulk phases. Many important phenomena, noticeable amongst these being corrosion, electrode processes, heterogeneous catalysis, dissolution and crystallisation occur at interfaces. The subject of surface chemistry finds

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